
Numericals Chemistry Chapter Solid State

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Numericals Chemistry **OMB No.**
Chapter Solid State **9753160843745 edited**
 by

VICTORIA CAROLYN

Numericals Chemistry Chapter Solid State
What type of solid is it? Ans: It is a covalent or network solid.
1.8. Why are ionic solids conducting in the molten state and not in the solid state? Ans: In the ionic solids, the electrical conductivity is due to the movement of the ions. Since the ionic mobility is negligible in the solid state, these are non-conducting in this state.
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Solids Solids are the chemical substances which are characterised by define shape and volume, rigidity, high density, low compressibility. The constituent particles

(atoms, molecules or ions) are closely packed and held together by strong interparticle forces
Types of Solids
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1. Density of unit cell: $d = \frac{z \cdot M}{a^3 \cdot N_A}$ where, a is edge of unit cell
 $N_A =$ Avogadro number (6.023×10^{23})
 $M =$ Molar mass
 $z =$ number of atoms per unit cell
For fcc, $z =$

4 for bcc, $z = 2$ for simple cubic, $z = 1$
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subject chemistry by testing themselves
on different and unique questions. Solid
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Chemistry The Solid State PREVIOUS
YEARS' QUESTIONS 2015 Very Short
Answer Type Questions [1 Mark] 1. What is
the formula of a compound in which the
element Y forms ccp lattice and atoms of X
occupy $1/3$ rd of tetrahedral voids? Ans. 2.
What is the formula of a compound
[...] Important Questions for CBSE Class 12
Chemistry The Solid ... Question: 1.7 - Solid
A is a very hard electrical insulator in solid
as well as in molten state and melts at

extremely high temperature. What type of solid is it? Answer: Given solid 'A' is a covalent solids, such as diamond.

Question: 1.8 - Ionic solids conduct electricity in molten state but not in solid state. Explain.

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What type of solid is it? Ans: It is a covalent or network solid.

1.8. Why are ionic solids conducting in the molten state and not in the solid state ? Ans: In the ionic solids, the electrical conductivity is due to the movement of the ions. Since the ionic mobility is negligible in the solid state, these are non-conducting in this state.

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Solid state Formulae 1. Density of unit cell:

$d = \frac{Z \cdot M}{a^3 \cdot N_A}$ where, a is edge of unit

cell N_A = Avogadro number (6.023×10

23) M = Molar mass z = number of atoms

per unit cell For fcc, $z = 4$ for bcc, $z = 2$ for

simple cubic, $z = 1$.

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The Solid State Solids Solids are the

chemical substances which are

characterised by definite shape and

volume, rigidity, high density, low

compressibility. The constituent particles

(atoms, molecules or ions) are closely

packed and held together by strong

interparticle forces Types of Solids

Numericals Chemistry Chapter Solid State Solids have definite volume, shape, and mass due to the short distance between the fixed position of particles. To learn more access Solid State Chemistry Class 12 Notes PDF on BYJU'S.

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Question: 1.7 - Solid A is a very hard electrical insulator in solid as well as in molten state and melts at extremely high temperature. What type of solid is it?

Answer: Given solid 'A' is a covalent solids, such as diamond. Question: 1.8 - Ionic solids conduct electricity in molten state but not in solid state. Explain.

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