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principle can be applied to
understand the effect of
change in pressure on the
systems at equilibrium as
follows. 1) When the
partial pressure of any of
the gaseous reactants or

of the products is
increased , the position of
equilibrium is shifted so
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establish a new
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state Le Chatelier's
principle when a physical
or chemical system at
equilibrium is disturbed
by applying a stress
(changing the
temperature, pressure or
concentration), it acquires
a new equilibrium so as to
relieve the
stress. Chemfax
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to changes in pressure,
temperature, or
concentration. Le
Chatelier's principle
implies that the addition
of heat to a reaction will
favor the endothermic
direction of a reaction as
this reduces the amount
of heat produced in the
system.

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Le Chatelier's principle

can be used to predict the behavior of a system due to changes in pressure, temperature, or concentration. Le Chatelier's principle implies that the addition of heat to a reaction will favor the endothermic direction of a reaction as this reduces the amount of heat produced in the system.

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