

# Ap Chemistry Chapter 3 Test

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OMB No. 0293248834071 edited by

## HERRERA MCKENZIE

Ap Chemistry Chapter 3 Test AP Chemistry Chapter 3 Test AP Chemistry - Practice Test: Chapter 3 Name \_\_\_\_\_ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) Silver has two naturally occurring isotopes with the following isotopic masses: 107.47 Ar 107.47 Ar 106.90509 108.9047 The average atomic mass of silver is 107.8682 amu. AP Chemistry - Practice Test: Chapter 3 Ar AP Chemistry Test (Chapter 3) (November) Multiple Choice and FIB (40%) 1) A chemistry student is filtering and drying a precipitate that formed from two solutions reacting. Which one is the most likely reason for an exceptionally high % yield? A) Careless filtering and insufficient drying B) Careless filtering and sufficient drying AP Chemistry Test (Chapter 3) - Denton ISD AP Chemistry-Chapter 3 MC Practice Questions Multiple Choice Identify the choice that best completes the statement or answers the question. \_\_\_\_ 1. Combustion analysis of 0.800 grams of an unknown hydrocarbon yields 2.613 g CO<sub>2</sub> and 0.778 g H<sub>2</sub>O. What is the percent composition of the hydrocarbon? a. 66.6% C; 33.4% H AP Chemistry-Chapter 3 MC Practice Questions AP Chemistry Practice Test #1 - Key Chapters 1, 2, and 3 1. In 1928, 3.3 g of a new element was isolated from 660 kg of the ore molybdenite. The percent by mass of this element in the ore was: a. 5 % b. 6.6 % c. 3.3 % d. 5.0x10<sup>-4</sup> % e. 2.2 % 2. A quantitative observation a. contains a number and a unit. b. does not contain

a number. AP Chemistry Practice Test #1 - Chapters 1, 2, and 3 Phosphoric acid can be prepared by reaction of sulfuric acid with "phosphate rock" according to the equation: Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + 3H<sub>2</sub>SO<sub>4</sub> -> 3CaSO<sub>4</sub> + 2H<sub>3</sub>PO<sub>4</sub> Suppose the reaction is carried out starting with 103 g of Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> and 75.0 g of H<sub>2</sub>SO<sub>4</sub>. Which substance is the limiting reactant? AP Chemistry Chapter 3 Quiz - ProProfs Quiz AP Chemistry: Chapter 3 - Stoichiometry - Read online for free. Chapter 3: Stoichiometry topics covered- atomic masses, the mole, coefficients, balanced equations, mole ratio, mass relationships, empirical formula, molecular formula, atomic and molecular weights, & molar mass I made this review for the kids that I tutor, but I decided to post it on the internet so that other people can use it too AP Chemistry: Chapter 3 - Stoichiometry | Mole (Unit ... Learn chapter 3 test ap chemistry zumdahl with free interactive flashcards. Choose from 500 different sets of chapter 3 test ap chemistry zumdahl flashcards on Quizlet. chapter 3 test ap chemistry zumdahl Flashcards and Study ... AP Chemistry. Chapters 1, 2, and 3: Stoichiometry and Atomic Theory. ... Chapter 3: Stoichiometry: Calculations with Chemical Formulas and Equations. How can I study for my first Unit Test in AP Chemistry!!!! Relax; Look over the Study Guide; Review Mastering in Chemistry assignments; Mr. Scaringi's Chemistry Page - Chapters 1, 2, and 3 ... These are the answers and explanations to the practice test on Chapters 1 - 3, which can be found here: <https://goo.gl/NgVq75>. ... General Chemistry 1 Review Study Guide - IB, AP, ... Chapters 1 - 3 Practice Test AP Chemistry Practice Test: Chapter 4 Name \_\_\_\_\_ SHORT ANSWER.

Write the word or phrase that best completes each statement or answers the question. 1)What is the solvent in an aqueous solution? ESSAY. Write your answer in the space provided or on a separate sheet of paper. 2)Explain why water is a good solvent for ionic substances. ...AP Chemistry Practice Test: Chapter 4 SHORT ANSWER. Write ...AP Chemistry Test (Chapter 3) (Autumn) Multiple Choice and FIB (45%) 1) Which of these lab errors would cause a low % yield O 2? Δ 2 KClO 3 (s) 2 KCl (s) + 3 O 2 (g) I) Using a crucible that has soot stuck to it II) Using a low mass of KClO 3 III) Heating the crucible an insufficient length of time IV) Spilling some of the KCl

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Phosphoric acid can be prepared by reaction of sulfuric acid with “phosphate rock” according to the equation: Ca 3 (PO 4) 2 + 3 H 2 SO 4 -> 3 CaSO 4 + 2 H 3 PO 4 Suppose the reaction is carried out

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