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answers. Acid Formula of acid Base
 Formula of base AMOUNT OF
 SUBSTANCE the value of $-114.2 \text{ kJ mol}^{-1}$
 in the first equation means that 114.2 kJ
 of heat energy is released when 1 mole
 of H_2SO_4 reacts with 2 moles of NaOH .
 the value of $-57.1 \text{ kJ mol}^{-1}$ in the second
 equation means that 57.1 kJ of heat
 energy is released when $\frac{1}{2}$ mole of H_2SO_4
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 amount of energy needed to make 1 g of
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 the specific heat capacity (measured in $\text{J g}^{-1} \text{ K}^{-1}$). • The following equation is
 then used to find the amount of heat
 energy give out (or
 absorbed). Chemsheets AS 029
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- The amount of energy needed to make 1 g of a substance 1 °C (1 K) hotter is called the specific heat capacity (measured in $\text{J g}^{-1} \text{K}^{-1}$).
- The following equation is then used to find the amount of heat energy give out (or absorbed).

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1) Use your knowledge of ionic equations to give the molar ratio in which the following acids react with bases. Complete the table to show your answers. Acid Formula of acid Base Formula of base

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the value of $-114.2 \text{ kJ mol}^{-1}$ in the first equation means that 114.2 kJ of heat energy is released when 1 mole of H_2SO_4 reacts with 2 moles of NaOH . the value of $-57.1 \text{ kJ mol}^{-1}$ in the second equation means that 57.1 kJ of heat energy is released when $\frac{1}{2} \text{ mole}$ of H_2SO_4 reacts with 1 mole of NaOH .

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